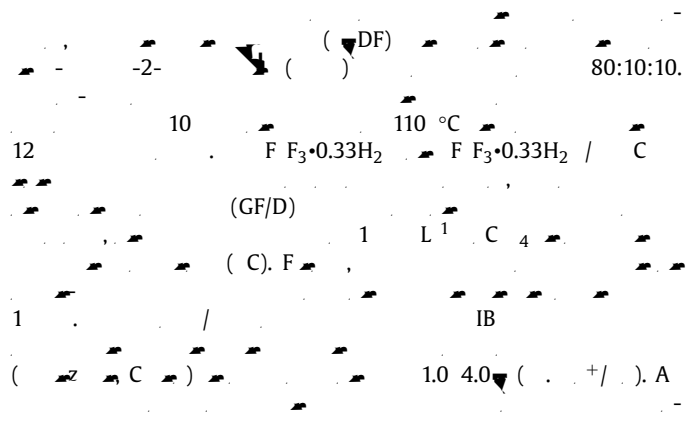
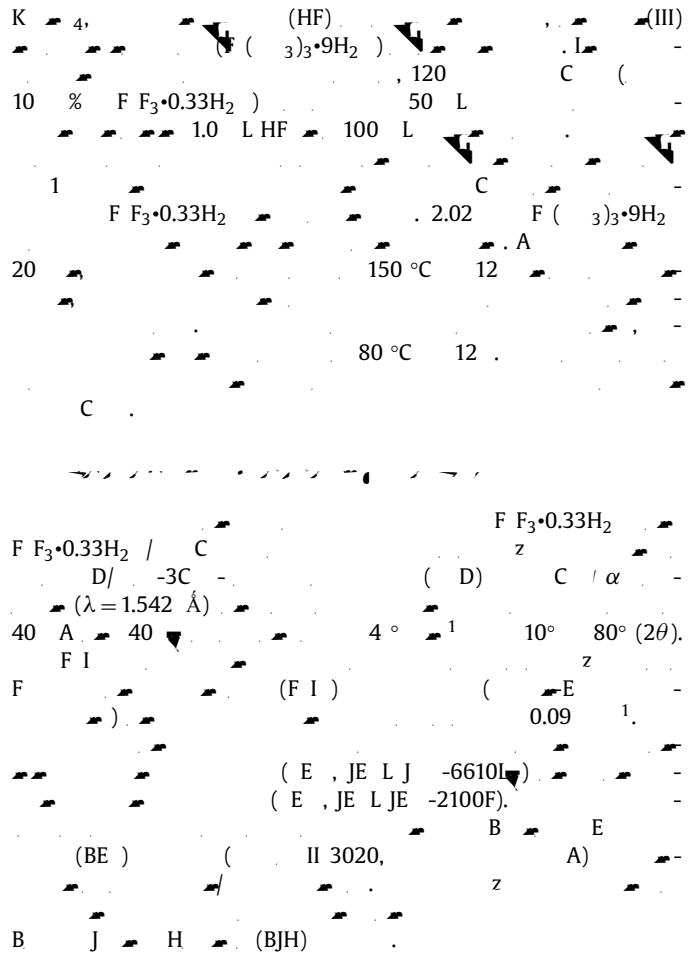
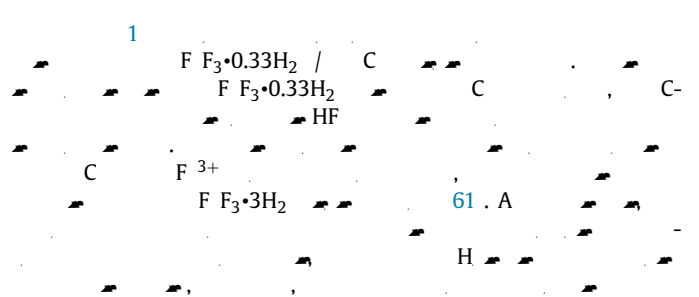


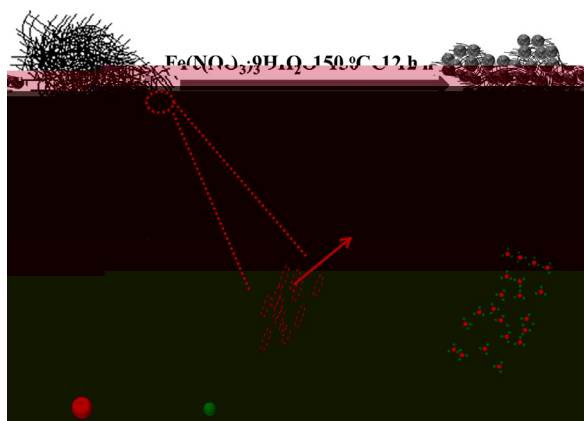
2. Experimental

$F F_3 \cdot 0.33H_2 / C$
 C
 H_2 4



3. Results and discussion





Scheme 1.

$F_3 \cdot 0.33H_2 / C$
 62 64
 F^{3+}
 A
 $F_3 \cdot 3H_2$
 27
 $F_3 \cdot 0.33H_2$
 24
 $F_3 \cdot 0.33H_2$
 C
 D
 $F_3 \cdot 1()$ I
 30°
 $2\theta = 13.8^\circ, 23.6^\circ, 27.8^\circ$
 $(1\ 1\ 0), (0\ 0\ 2), (2\ 2\ 0)$
 $F_3 \cdot 0.33H_2$ (JC D 76-1262).
 (H B) $F_3 \cdot 0.33H_2$ ($F_3 \cdot 1()$)
 B
 $H\ B-F_3 \cdot 0.33H_2$
 C
 $F_3 \cdot 0.33H_2 / C$
 $F_3 \cdot 2()$ 1343 1580 1

(D), A
 1612 1 (G), D'
 66. $F_3 \cdot 0.33H_2 / C$
 (1.80) I_D / I_G C (1.62) I_D / I_G
 $F_3 \cdot 0.33H_2$
 C 67. $F_3 \cdot 0.33H_2 / C$ ($F_3 \cdot 2()$)
 $F_3 \cdot 0.33H_2$
 C $F_3 \cdot 0.33H_2 / C$
 $F_3 \cdot 0.33H_2$
 H H
 530 1
 $F_3 \cdot 0.33H_2$
 1111 1 $F_3 \cdot 0.33H_2 / C$
 $F_3 \cdot 0.33H_2 / C$
 68. $F_3 \cdot 0.33H_2 / C$
 C

z E . A $F_3 \cdot 3()$ C
 $F_3 \cdot 3()$ $F_3 \cdot 0.33H_2$ A
 z 800 . B
 $F_3 \cdot 0.33H_2$
 z
 $F_3 \cdot$

60.
 C

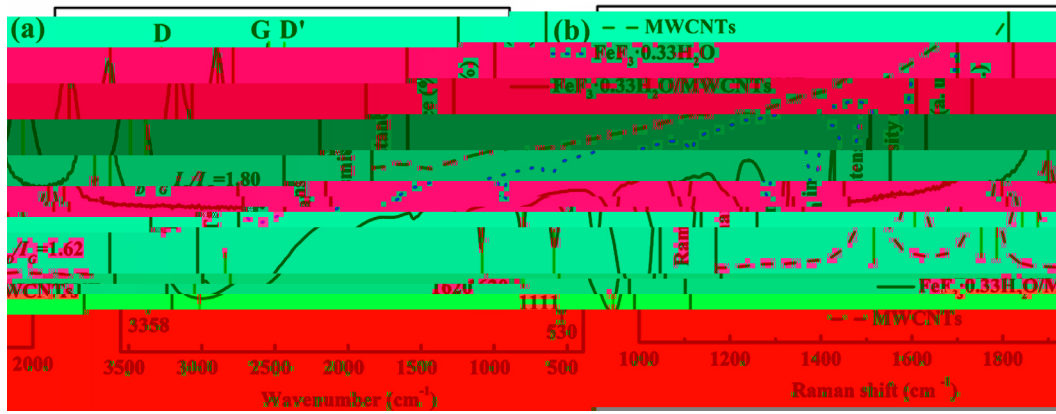


Fig. 2. (a) Raman spectra of FeF₃·0.33H₂O and FeF₃·0.33H₂O/MWCNTs. (b) Raman spectra of FeF₃·0.33H₂O and FeF₃·0.33H₂O/MWCNTs.

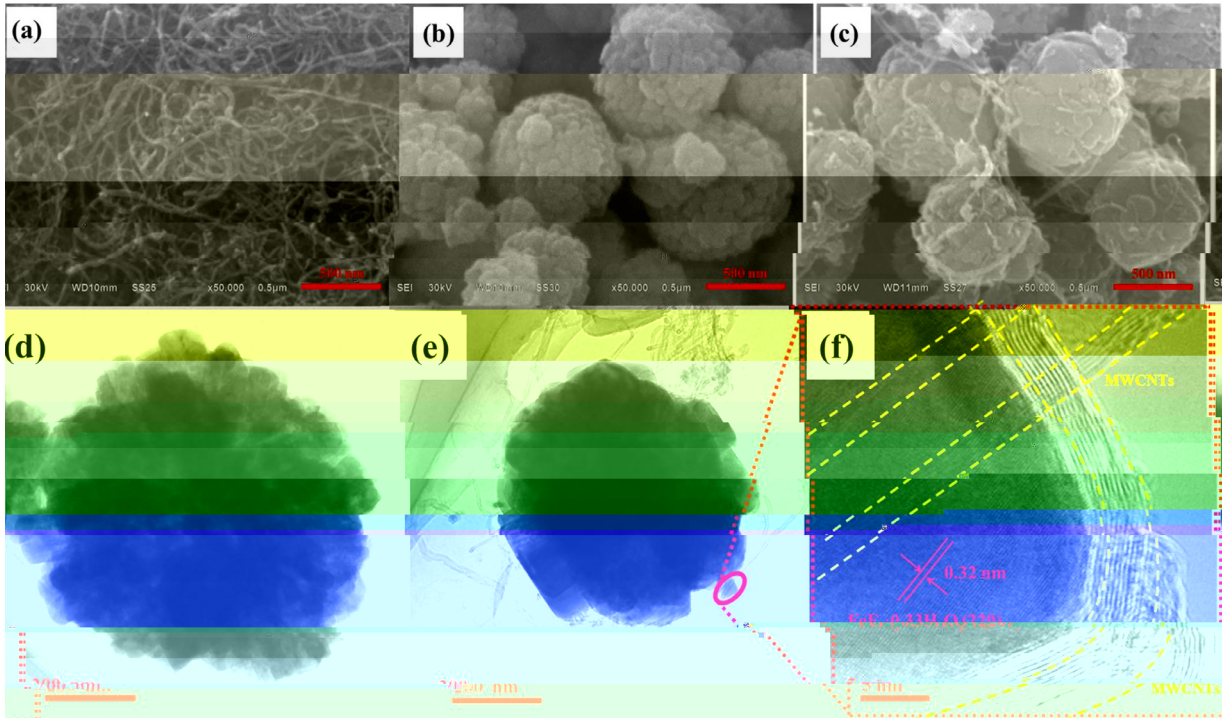
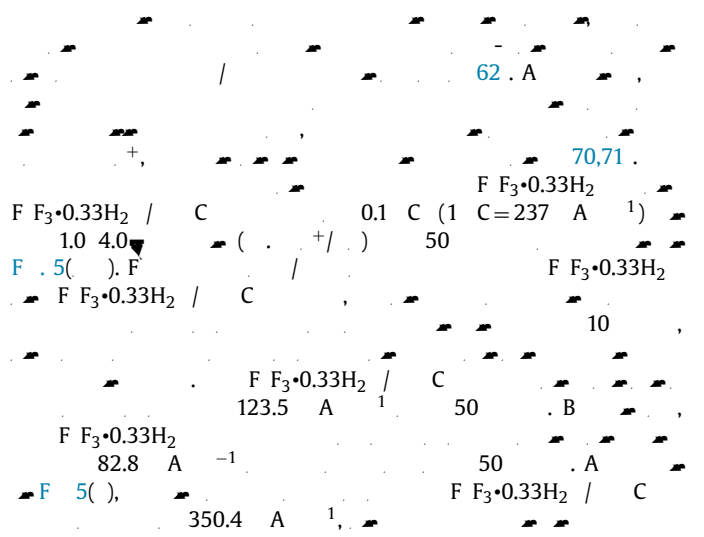
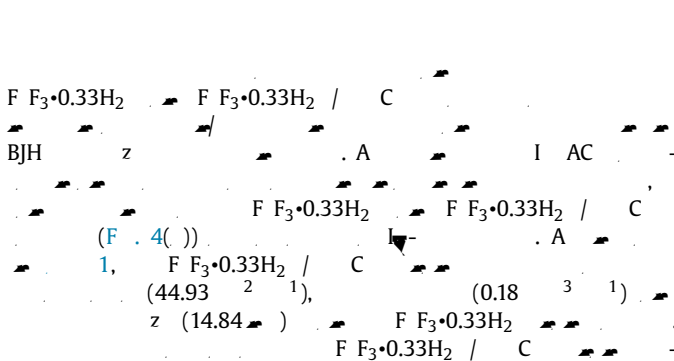


Fig. 3. (a) SEM image of FeF₃·0.33H₂O. (b) SEM image of FeF₃·0.33H₂O/MWCNTs. (c) SEM image of FeF₃·0.33H₂O/MWCNTs. (d) TEM image of FeF₃·0.33H₂O. (e) TEM image of FeF₃·0.33H₂O/MWCNTs. (f) HRTEM image of FeF₃·0.33H₂O/MWCNTs.

	FeF ₃ ·0.33H ₂ O	FeF ₃ ·0.33H ₂ O/MWCNTs
BE		
(² 1)		
BJH		
(³ 1)		
z (μm)		
FeF ₃ ·0.33H ₂ O	33	0.067
FeF ₃ ·0.33H ₂ O/MWCNTs	45	0.18
		14.9



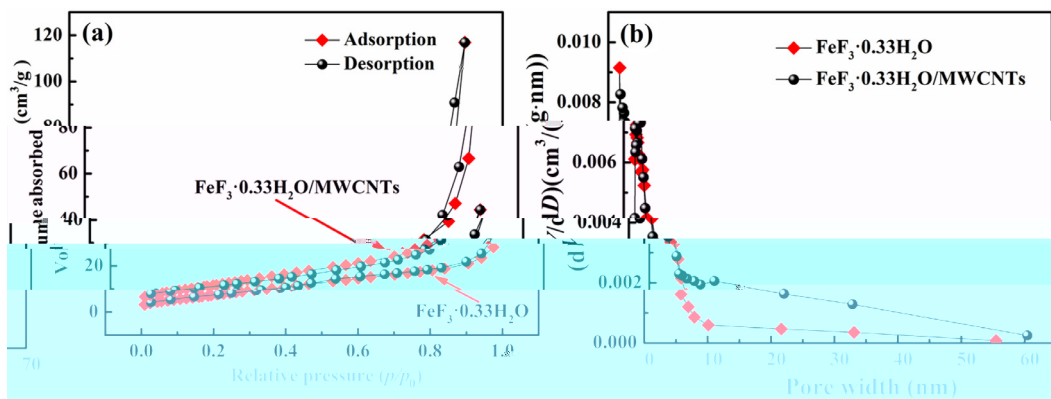


Fig. 4. (a) Nitrogen adsorption-desorption isotherm of FeF₃·0.33H₂O/MWCNTs (red diamonds) and FeF₃·0.33H₂O (black circles). (b) Pore size distribution of FeF₃·0.33H₂O/MWCNTs (red diamonds) and FeF₃·0.33H₂O (black circles) calculated by BJH method.

Sample	Surface Area (A)	Volume (V)	Method
FeF ₃ ·0.33H ₂ O / C	294.4	1	A
FeF ₃ ·0.33H ₂ O / C	246.0	5	A
FeF ₃ ·0.33H ₂ O / C	211.6	10	A
FeF ₃ ·0.33H ₂ O / C	163.8	30	A

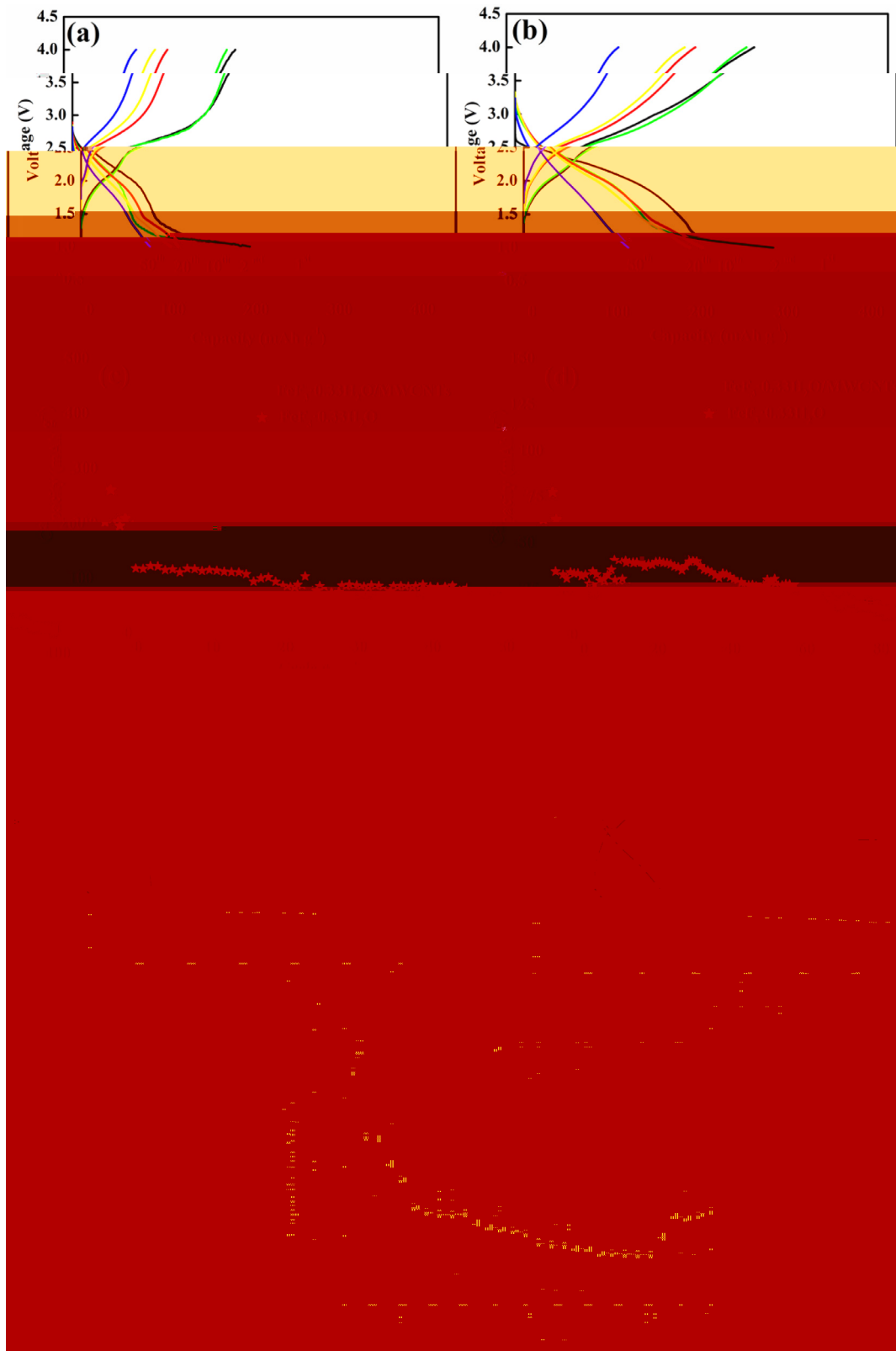


Fig. 5. (a) Cyclic voltammograms of Pt/C electrode in 0.1 M $F_3 \cdot 0.33H_2SO_4$ solution at various scan rates (1, 2, 10, 20, 50 mV/s); (b) Cyclic voltammograms of Pt/C electrode in 0.1 M $F_3 \cdot 0.33H_2SO_4$ solution at scan rates of 1, 2, 10, 20, 50, and 100 mV/s ; (c) Nyquist plot of Pt/C electrode in 0.1 M $F_3 \cdot 0.33H_2SO_4$ solution at a scan rate of 100 mV/s . The inset shows the equivalent circuit model for the Pt/C electrode.

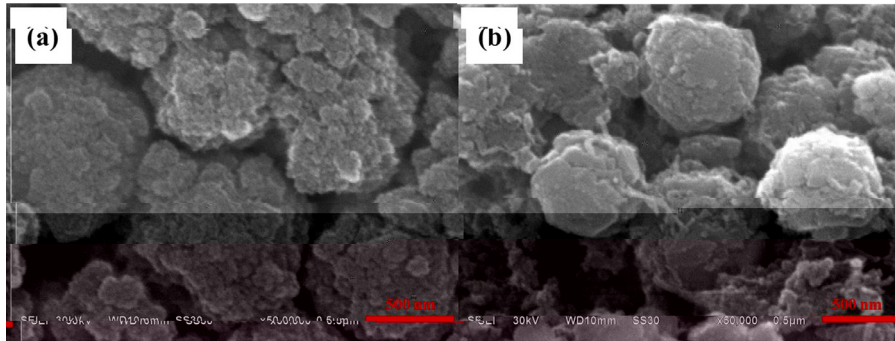


Fig. 6. EDS spectra of (a) F₃*0.33H₂ and (b) F₃*0.33H₂ / C at 50 and 0.1 C.

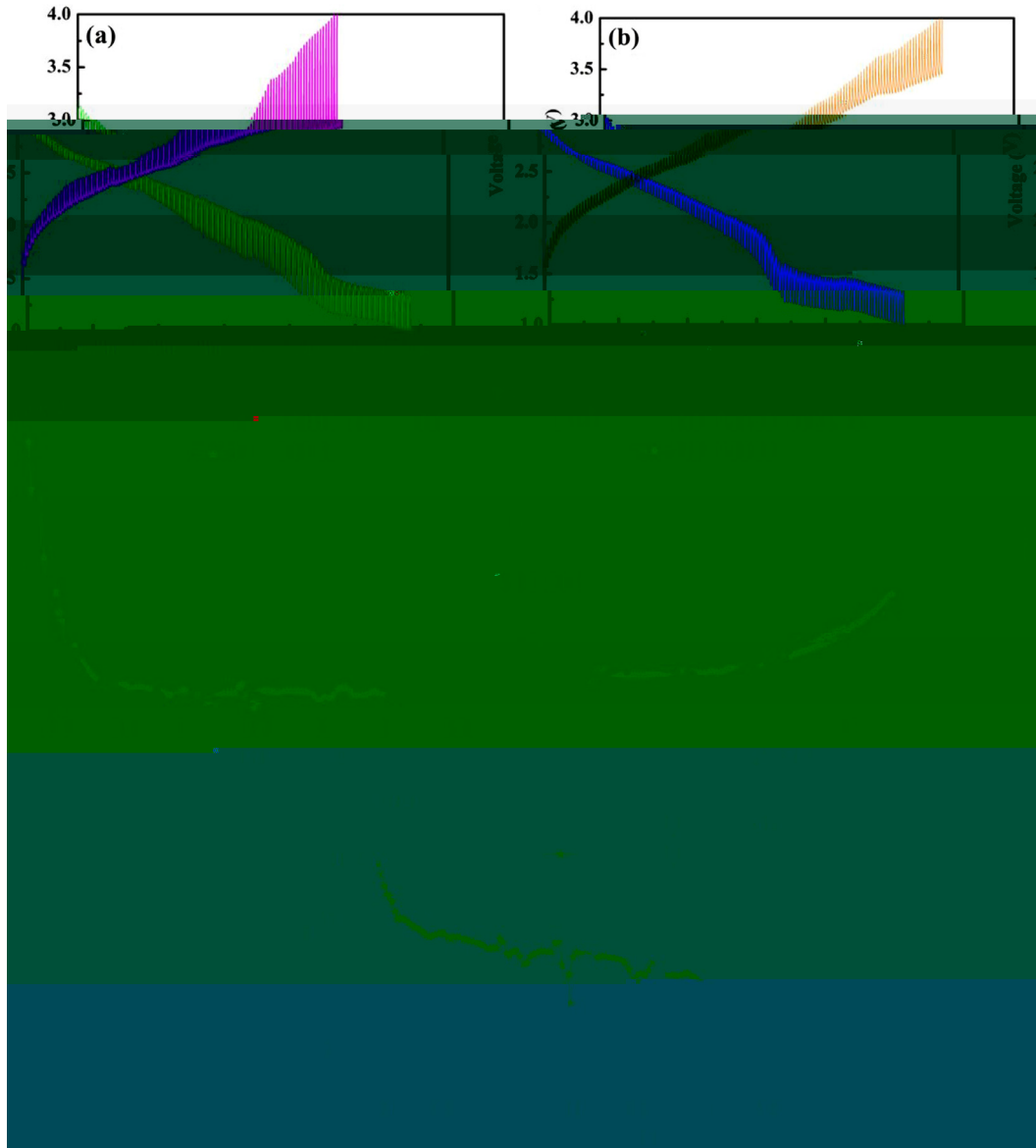
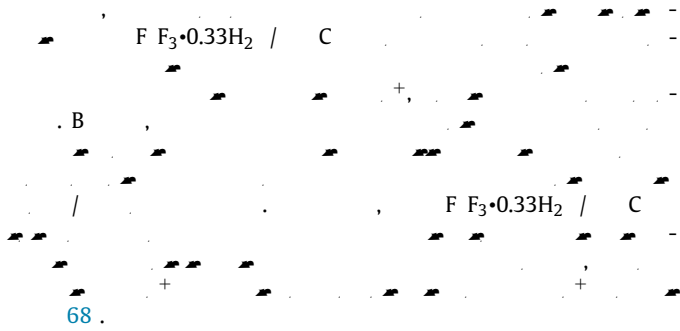
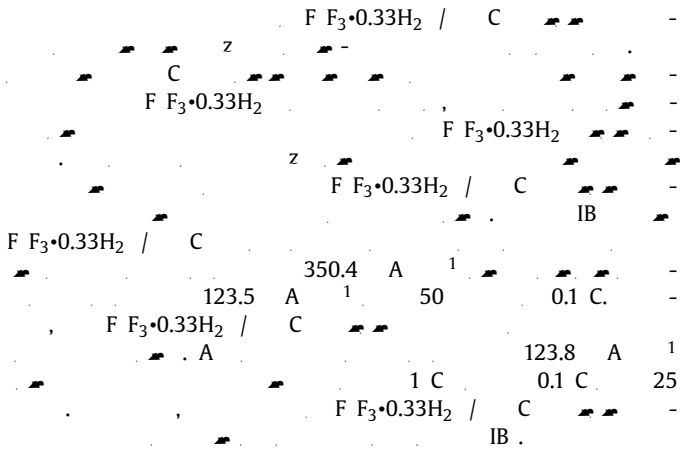


Fig. 7. CV curves and EIS Nyquist plots of (a) F₃*0.33H₂ and (b) F₃*0.33H₂ / C at various scan rates. The inset shows the equivalent circuit model. The parameters are: R_s = 1.0 Ω, R_{ct} = 4.0 Ω, CPE = 1.0 s^{1/2}. The values of R_{ct} and CPE are 2.7 × 10¹⁴ Ω and 6.5 × 10¹⁰ s^{1/2}, respectively.

$$Z_{CPE} = \frac{1}{j\omega CPE} \approx \frac{1}{j\omega} \ll \frac{1}{\omega} \quad (3)$$



4. Conclusions



Acknowledgments

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Supplementary materials

Supplementary materials are available for this article. For more information, please go to the journal web site at <http://dx.doi.org/10.1016/j.procs.2017.10.032>.

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63 F. . . , H. H. , H.B. . . . , G. . . , .C. . .